General Science

CH- 6, Physical and chemical changes

Name: _____ Date:_____ Class: VII Sec: ___

I. OBJECTIVE TYPE QUESTIONS:

1. What kind of change is the explosion of fireworks?

a) Physical change b) Chemical change

c) Not a change d) Reversible change

2. Which of the following is not a required condition for rusting to occur?

a) Presence of oxygen b) Presence of carbon dioxide

c) Presence of water vapour d) Presence of water

3. Which of the following describes the process of depositing a zinc layer on iron?

a) Galvanisation b) Crystallisation

c) Rusting d) None of the above

4. Which of the following chemicals is responsible for making lime water milky?

a) Calcium carbonate b) Calcium hydroxide

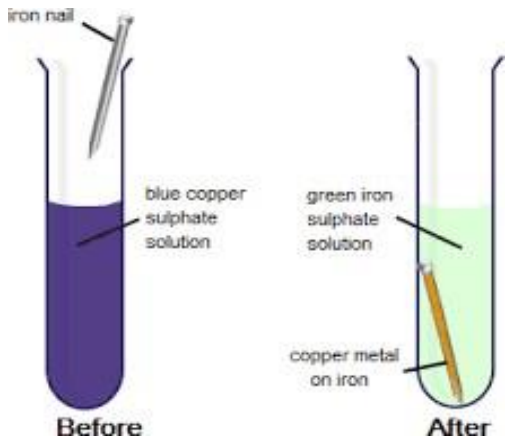
c) Calcium chloride d) Calcium oxide

5. When you leave a piece of iron in the open for some months, it acquires a film of brownish substance. The substance is called-

a) Acid b) Base

c) Rust d) None of these

6.	Which of the following is not a chemical change- a) Digestion of food b) Photosynthesis c) Burning of coal d) Melting of ice
II.	**Answer the following in short.**
1.	Give two examples for each of the following cases: a. Physical changes which are reversible. _____ b. Physical changes which are not reversible. _____
2.	What are the components of stainless steel? _____
3.	What is galvanization? Give its advantage. _____ _____
4.	Write any three methods to prevent iron from rusting. _____ _____
III.	**Fill in the blanks**
1.	Copper sulphate is also known as _____**.**
2.	Iron sulphate is also known as _____**.**
3.	Baking soda is _____
4.	_____ layer protects us from harmful ultraviolet radiation.
5.	_____ is a reddish brown, soft, porous, flaky coating on iron articles.
IV.	**Write the word equation for the following**
1.	Burning of magnesium _____

2.	Rusting of iron
3.	_____ Reaction between iron nail and copper sulphate
V.	**Complete the following**
	1. Magnesium oxide + water → _____ 2. Vinegar + baking soda → _____ + _____ **3.** CO_2 + lime water → _____ + _____
VI	**Observe the given picture and answer the questions.**  **a)** What change will you observe in the colour of the solution after dropping an iron nail into it? _____ **b)** Why do we observe the colour change in the solution? _____ **c)** What causes brown deposition on the iron nail? _____ **d)** Write the word equation involved in the above reaction. _____ → _____